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**Chapter 11 Review Molecular Composition Of Gases Section**

CHAPTER 11 REVIEW Molecular Composition of Gases MIXED REVIEW SHORT ANSWER Answer the following questions in the space provided. 1. The average speed of a gas molecule is most directly related to the . (a) polarity of the molecule (b) pressure of the gas (c) temperature of the gas (d) number of moles in the sample 2.

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CHAPTER 11 REVIEW Gases SECTION 4 SHORT ANSWER Answer the following questions in the space provided. 1. \_\_\_\_ List the following gases in order of increasing rate of effusion. (Assume all gases are at the same temperature and pressure.) (a) He (b) Xe (c) HCl (d) Cl<sub>2</sub> 2.

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molecular ones ∅ The intermolecular forces of the remaining substances depend on molecular weight, polarity, and hydrogen bonding. The molecular weights are H<sub>2</sub> (2), CO (28), HF (20), and Ne (20). ∅ The boiling point of H<sub>2</sub> should be the lowest because it is nonpolar and has the lowest molecular weight.

### **Chapter 11 Intermolecular Forces, Liquids, and Solids**

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According to its molecular formula, each molecule of sucrose contains 12 carbon atoms, 22 hydrogen atoms, and 11 oxygen atoms. A mole of sucrose molecules therefore contains 12 mol of carbon atoms, 22 mol of hydrogen atoms, and 11 mol of oxygen atoms.

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1. Describe Mendel's scientific approach and how it helped make his experiments successful. 2. Explain why the garden pea was a good choice for Mendel's experiments. 3. Contrast blending and the particulate concept of inheritance. 4.

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CHM 2046 Test 1 Review: Chapter 11, Chapter 12, & Chapter 13 Chapter 11: \*Below is a phase diagram for carbon dioxide. Use it to answer questions 1 and 2.\*

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CHAPTER 10 REVIEW States of Matter SECTION 3 SHORT ANSWER Answer the following questions in the space provided. 1. Match description on the right to the correct crystal type on the left. b ionic crystal (a) has mobile electrons in the crystal c covalent molecular crystal (b) is hard, brittle, and nonconducting

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Chapter 11 - Molecular Composition of Gases 11-1 Volume-Mass Relationships of Gases I. Measuring and Comparing the Volumes of Reacting Gases A. Observations of Gay-Lussac 1. 2 liters H<sub>2</sub> + 1 liter O<sub>2</sub> → 2 liters H<sub>2</sub>O vapor 2. 2 volumes H<sub>2</sub> + 1 volume O<sub>2</sub> → 2 volumes H<sub>2</sub>O vapor 3. 1 volume H<sub>2</sub> + 1 volume Cl<sub>2</sub> → 2 volumes HCl 4.

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McDougal Littell Geometry, Larson et al., McDougal Littell, 2004. Introduction. ... 2ND SEMESTER REVIEW AND EXAM N/A 9 3 Geometry A/H. Charlottesville City Schools.

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