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OBJ: 9.1.3 Identify the two common endings for the names of most polyatomic ions. | 9.5.3 Apply the rules for writing chemical formulas by using a flowchart. STA: Ch.2 33.ANS: OBJ: 9.5.3 Apply the rules for writing chemical formulas by using a flowchart.

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When the anion ends in -ite, the acid name is the stem of the anion with the suffix -ous. Example H<sub>2</sub>SO<sub>3</sub> (X=Sulfite), Sulfurous acid When the anion name ends in -ate, the acid name is the stem of the anion with the suffix -ic.

### **Chapter 9: Chemical Names and Formulas**

Names and Formulas for Bases A base is an ionic compound that produces hydroxide ions (OH<sup>-</sup>) when dissolved in water (the Arrhenius definition) Bases are named the same way as other ionic compounds: "The name of the cation (which is a metal) is followed by the name of the anion (which will be hydroxide).

### **Chapter 9 Chemical Names and Formulas**

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### **Chemistry Chapter 9 Chemical Names Formulas Answers**

Chapter 9 Chemical Names and Formulas 83 SECTION 9.3 NAMING AND WRITING FORMULAS FOR MOLECULAR COMPOUNDS (pages 268–270) This section explains the rules for naming and writing formulas for binary molecular compounds. Naming Binary Molecular Compounds (pages 268–269) 1. Circle

the letter of the type(s) of elements that form binary molecular compounds.

### **Name Date Class CHEMICAL NAMES AND FORMULAS 9**

282 Chapter 9 Chemical Reactions Section 9.1.1 Figure 9.1 When adipoyl chloride in dichloromethane reacts with hexanediamine, nylon is formed. Nylon is used in many products, including carpeting, clothing, sports equipment, and tires. Objectives Recognize evidence of chemical change. Represent chemical reactions with equations.

#### **Chapter 9: Chemical Reactions**

\_\_\_\_ 18. What is the name of  $\text{H}_2\text{SO}_3$ ? a. hyposulfuric acid c. sulfuric acid b. hydrosulfuric acid d. sulfurous acid \_\_\_\_ 19. When the name of an anion that is part of an acid ends in -ite, the acid name includes the suffix \_\_\_\_\_. a. -ous c. -ate b. -ic d. -ite \_\_\_\_ 20. What is the formula for sulfurous acid? a.  $\text{H}_2\text{SO}_4$  c.  $\text{H}_2\text{SO}_2$  b.  $\text{H}_2\text{SO}_3$  d.  $\text{H}_2\text{S}$  \_\_\_\_ 21.

#### **Chapter 9 Practice Test**

â€¢ The names of most polyatomic anions end in -ite or -ate. 9.2 Naming and Writing Formulas for Ionic Compounds â€¢ The name of a binary ionic compound is the cation name followed by the anion name.

#### **Chapter-09---Chemical-Names-and-Formulas.pdf - CHAPTER**

The name of a base is the name of the cation followed by the name of the anion (hydroxide). The formula of a base is written by showing the number of hydroxide ions needed to balance the positive charge on the cation. After reading Lesson 9.4, answer the following questions. Names and Formulas of Acids hydrogen1.

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Chapter 9 Chemical Names and Formulas 221 . 4. Write formulas for the ionic compounds. e. onium chloride f. potassium dichromate a. sodium phosphate b. magnesium sulfate c. sodium hydroxide d. potassium cyanide 5. Write formulas for compounds formed from these pairs of ions. a.  $\text{NH}_4^+$ .  $\text{SO}_4^{2-}$

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2. Biological, chemical, and physical properties of matter result from the ability of atoms to form bonds from ... a. Name the elements in the order listed in the formula. ... Chapter 9.PDF Author: tamara Created Date:

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Convert between the names of compounds and their chemical formulas. (Section 5.3) Report the answers to calculations to the correct number of significant figures. (Section 8.2) Use percentages as conversion factors. ... 332 Chapter 9 Chemical Calculations and Chemical Formulas.

## Chapter 9 Chemical Names and Formulas - Mark Bishop

Chapter 9 Chemical Names And Formulas Pdf name date class chemical names and formulas 9 - chapter 9 chemical names and formulas85 names and formulas for bases (page 273) 6. a base is a compound that produces \_\_\_\_\_ when

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II. Chemical formulas A. chemical formula 1) symbols representing the composition of the smallest unit of a substance 2) shows which elements are present and how many there are  $H_2SO_4$  = hydrogen, sulfur, oxygen = 7 atoms total B. "molecular formula symbols representing the composition of a molecular compound

## Ch. 5 Notes: TYPES OF COMPOUNDS - Kwanga.net

name date class chemical names and formulas 9 - chapter 9 chemical names and formulas85 names and formulas for bases (page 273) 6. a base is a compound that produces \_\_\_\_\_ when dissolved in water.

## Chapter 9 Chemical Names And Formulas Answers PDF

Chapter 9 "Chemical Names and Formulas" • Click to add text  $H_2O$  2 Section 9.1 Naming Ions OBJECTIVES: "Identify the charges on monatomic ions by using the periodic table, and name the ions. 3 Section 9.1 Naming Ions OBJECTIVES: "Define a polyatomic ion and write the names and formulas of the most common polyatomic ions. 4 Section 9.1 ...

## Chapter 09 Chemical Names and Formulas (handout)

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## Chapter 9: Cellular Respiration: Harvesting Chemical Energy

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## Prentice Hall Chemistry Chapter 9: Chemical Names and

Law of Definite Proportions- in a sample of a chemical compound, the masses of the elements are always in the same proportions.  $H_2O$  (water) and  $H_2O_2$  (hydrogen peroxide) Some Laws: 2. Law of Multiple

Proportions- Dalton stated that whenever two elements form more than one compound, the ... Chapter 9 Chemical Names and Formulas Last ...

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### **9.1 Naming Ions 9 - Ms. Lara La Cueva HS Science**

Chapter 9 Chemical Names and Formulas. 233 \_\_\_\_\_ 12. Nitrogen reacts with oxygen to form two compounds. Compound A contains 2.8 g of nitrogen for each 1.6 g of oxygen.

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### **CHAPTER 9 CHEMICAL NAMES AND FORMULAS ANSWERS PDF**

29. Given the name of a binary molecular compound, describe how to write its formula. Use carbon tetrachloride as an example. 30. Name the compounds  $\text{CuBr}_2$ ,  $\text{SCl}_2$ , and  $\text{BaF}_2$ . Explain the use or omission of the Roman numeral (II) and the prefix di-

### **Chemical Names and Formulas Test - Mr. Walk**

Chapter 9 Ź“ Chemical Names and Formulas. Jennie L. Borders. Periodic Table for Naming. Section 9.1 Regular Metals Review. The regular metals include Group 1 (except H) and 2 and aluminum. When naming a compound that starts with a regular metal, name the first element and add Ź“ide to the second element (except for polyatomic ions).

### **Chapter 9 Ź“ Chemical Names and Formulas**

122 Chapter 9 Chemical Formulas The names of the iŹrst twoexamples correspond sensibly to the formulas

of the substances, that for acetylene is a structural formula (the molecular formula would be

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